

Step Academy official

Model Town Grw PH: 03016652757

STUDENT NAME	
PAPER CODE	77636
TIME ALLOWED	40
Paper Date	14-02-2026



CLASS	New 1st Year (FSC/ICS)
SUBJECT	Chemistry
TOTAL MARKS	25
Paper Type	

Q1. Choose the correct answer.

5X1=5

1. Which one do you think is correct about metallic character?

(A) It decreases from top to bottom in a group. (B) It increases from top to bottom in a group. (C) It remains constant from left to right in a period. (D) It increases from left to right in a period.

2. What is the oxidation state of sulfur in the sulfate ion (SO_4^{2-})?

(A) +4 (B) +2 (C) +6 (D) 0

3. Which one of the following elements has the highest ionization energy?

(A) Sodium (Na) (B) Magnesium (Mg) (C) Aluminium (Al) (D) Argon (Ar)

4. The law of triad applies to the following set of elements:

(A) B, N, C (B) Be, Mg, Ca (C) Ar, K, Ca (D) Te, As

5. The ionization energy of N is higher than that of oxygen because of:

(A) Greater attraction of electrons by the nucleus (B) The half-filled p-orbitals of N possess extra stability (C) Greater penetration effect (D) The size of the N atom being smaller

Q2. Write short answers of the following questions.

5X2=10

1. What is the importance of the development of the periodic table?

2. Give salient features of period 4 in the modern Periodic Table.

3. What is s-block elements and why are they called so?

4. Why E.A of oxygen atom is less than Sulphur in Group VI-A?

5. How does the nature of orbital influence the value of ionization energies of elements?

Q3. Write detailed answers of the following questions.

2X5=10

1. Explain with the help of equations, acidic and basic behavior of oxides and chlorides.

2. Describe the factors affecting and periodic trends of electron affinity.