

Step Academy official

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STUDENT NAME	
PAPER CODE	77636
TIME ALLOWED	40
Paper Date	14-02-2026



CLASS	New 1st Year (FSC/ICS)
SUBJECT	Chemistry
TOTAL MARKS	25
Paper Type	

Q1. Choose the correct answer.

5X1=5

1. Which one do you think is correct about metallic character?

- (A) It decreases from top to bottom in a group. (B) It increases from top to bottom in a group. (C) It remains constant from left to right in a period. (D) It increases from left to right in a period.

2. What is the oxidation state of sulfur in the sulfate ion (SO_4^{2-})?

- (A) +4 (B) +2 (C) +6 (D) 0

3. Which one of the following elements has the highest ionization energy?

- (A) Sodium (Na) (B) Magnesium (Mg) (C) Aluminium (Al) (D) Argon (Ar)

4. The law of triad applies to the following set of elements:

- (A) B, N, C (B) Be, Mg, Ca (C) Ar, K, Ca (D) Te, As

5. The ionization energy of N is higher than that of oxygen because of:

- (A) Greater attraction of electrons by the nucleus (B) The half-filled p-orbitals of N possess extra stability (C) Greater penetration effect (D) The size of the N atom being smaller

Q2. Write short answers of the following questions.

5X2=10

1 . What is the importance of the development of the periodic table?

2 . Give salient features of period 4 in the modern Periodic Table.

3 . What is s-block elements and why are they called so?

4 . Why E. A of oxygen atom is less than Sulphur in Group VI-A?

5 . How does the nature of orbital influence the value of ionization energies of elements?

Q3. Write detailed answers of the following questions.

2X5=10

1 . Explain with the help of equations, acidic and basic behavior of oxides and chlorides.

2 . Describe the factors affecting and periodic trends of electron affinity.